QP CODE:212006 (NEW SCHEME) **Reg.No: ……..**

Second Year B.Pharm Degree Supplementary Examinations - March 2015

(2012 Scheme)

PHARMACEUTICAL ANALYSIS

Time: 3 Hours **Total Marks: 100**

- Answer all Questions.
- Write equations wherever necessary.

Essay **(3x10=30)**

1. Explain the theoretical principles involved, equivalence point determination and titration curves of precipitation titrations
2. What is law of mass action. Describe any three of its important applications of chemical analysis
3. How do gravimetric precipitates are washed and ignited. Discuss the gravimetric estimation of magnesium.

Short notes **(14x5=70)**

4. Define hydrolysis of salts and solubility product with suitable examples
5. Name five important primary standards for various types of titrations and its equivalent weights & conditions of use
6. Thermo gravimetric curves
7. Explain the difference in iodimetric and iodometric titrations
8. Describe the necessary precautions required to be observed to avoid errors in titrimetric analysis
9. What is meant by standard deviation. How will you determine standard deviation in results obtained by summation of data
10. Explain the oxygen flask combustion method
11. Outline the theory involved in non aqueous titrations
12. How will you prepare and standardize 0.1 M lithium meth oxide
13. Explain the effect of acids, temperature and solvent on the solubility of a precipitate
14. What is meant by PH and PM. Explain it
15. Various methods employed for carrying out EDTA titrations.
16. Sodium 2, 6 – dichloro phenol indophenol titrations with suitable examples
17. Explain the method of assay of oxygen and carbon dioxide

****