QP CODE:202006 **Reg.No: ……..**

Second Year B.Pharm Degree Supplementary Examinations - April 2014

PHARMACEUTICAL ANALYSIS

Time: 3 Hours **Total Marks: 100**

- Answer all Questions.
- Write equations wherever necessary.

Essay **(3x10=30)**

1. Explain principle of precipitation titrations. Discuss the effect of temperature and solvent on the solubility of precipitate.
2. What do you understand by complexation and chelation. Explain briefly about the titration curves in complexometry.
3. What is the importance of non aqueous titration in pharmaceutical analysis and mention its limitations . Explain principle and procedure for the standardization of per chloric acid (0.1 M).

Short notes **(14x5=70)**

4. Discuss Arrhenius and Bronsted – Lowry concepts of acid-base.
5. Explain alkalimetry in non-aqueous system.
6. Explain Fajan's and Mohr's method.
7. Mention the difference between iodimetry and iodometry by giving suitable reactions.
8. Importance of buffer in complexometric titrations.
9. Polyprotic system.
10. Explain the assay of carbon dioxide using nitrometer.
11. What are the different methods of expressing concentrations.
12. Acid-base indicators.
13. Types of solvents used in non-aqueous titrations.
14. Mention the techniques of drying and ignition of the precipitate in gravimetric analysis.
15. Detection of end point in redox titration.
16. Explain the method of estimation of nitrogen.
17. Explain the preparation and standardization of 0.1 M sodium thiosulphate solution.

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